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## **Covalent Bonding And Molecular Structure**

The chemical bonds formed are so-called, covalent bonds. A covalent bond is formed between two nuclei so that resulting molecule is stabilized, hence

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existence of the molecule. The interactions of particles in H<sub>2</sub> molecule for example is only electromagnetic in nature, more specifically Coulomb interactions.

### **Chapter 7. Covalent Bonds and Molecular Structure**

- Describe the formation of single

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covalent bonds in  $H_2$ ,  $Cl_2$ ,  $H_2O$ ,  $CH_4$ ,  $NH_3$  and  $HCl$  as the sharing of pairs of electrons leading to the noble gas configuration. A covalent bond is formed when atoms share electrons. Eg 1: Hydrogen A hydrogen atom has only one electron in the first shell.

## **covalent-bonding-and-molecules**

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A covalent bond is a chemical bond that involves the sharing of electron pairs between atoms. These electron pairs are known as shared pairs or bonding pairs, and the stable balance of attractive and repulsive forces between atoms, when they share electrons, is known as covalent bonding. For many molecules, the sharing of electrons allows each



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atom to attain the equivalent of a full outer shell ...

## **Covalent bond - Wikipedia**

The term covalent molecular structure describes molecules having covalent bonds. A molecule is a group of atoms bonded together through chemical bonds. When these bonds are covalent

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bonds, these molecules are known as covalent molecular compounds. These covalent molecular structures can be either polar compounds or nonpolar compounds depending on the electronegativity of the atoms that are involved in bond formation.

## **Difference Between Covalent**

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### **Molecular and Covalent Network ...**

Ionic bonds result from the transfer of electrons from one atom to another (formed by a metal and a non-metal)

Covalent bonds result from two atoms sharing electrons (formed by 2 or more non-metals). 2) How are nonpolar

covalent bonds different from covalent bonds, and what types of elements

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combine to form each?

## **Chapter 12 Review 1: Covalent Bonds and Molecular Structure**

Covalent Bonding and Molecular  
Structure (key) AX 2: CO 2 1. What is the  
O-C-O bond angle?  $180^\circ$  2. Which  
element is more electronegative,  
carbon, or oxygen? Oxygen 3. Would

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you expect CO<sub>2</sub> to be a polar or a non-polar molecule? Explain. Polar. Even though it is composed of polar bonds, the two bonds are located opposite one

### **Covalent Bonding and Molecular Structure (key)**

As a general rule, covalent bonds are formed between elements lying toward

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the right in the periodic table (i.e., the nonmetals). Molecules of identical atoms, such as  $H_2$  and buckminsterfullerene ( $C_{60}$ ), are also held together by covalent bonds. Lewis formulation of a covalent bond. In Lewis terms a covalent bond is a shared electron pair.

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## **Chemical bonding - Covalent bonds | Britannica**

The hydrogen molecule provides a simple example of MO formation. In the following diagram, two 1s atomic orbitals combine to give a sigma ( $\sigma$ ) bonding (low energy) molecular orbital and a second higher energy MO referred to as an antibonding orbital. The

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bonding MO is occupied by two electrons of opposite spin, the result being a covalent bond.

### **Molecular Structure & Bonding**

An example is carbon dioxide ( $\text{CO}_2$ ), the molecules of which contain one atom of carbon bonded with two atoms of oxygen. However, although the covalent



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bonds holding the atoms together in a simple...

### **Simple covalent molecules - Structures and properties ...**

Covalent bonds involve the sharing of electron pairs between atoms. Electron pairs shared between atoms of equal or very similar electronegativity constitute

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a nonpolar covalent bond (e.g., H-H or C-H), while electrons shared between atoms of unequal electronegativity constitute a polar covalent bond (e.g., H-O).

## **Covalent bonds (video) | Chemistry of life | Khan Academy**

When a covalent bond is formed, the

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atomic orbitals (the orbitals in the individual atoms) merge to produce a new molecular orbital which contains the electron pair which creates the bond. Four molecular orbitals are formed, looking rather like the original  $sp^3$  hybrids, but with a hydrogen nucleus embedded in each lobe.

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## **COVALENT BONDING - SINGLE BONDS - chemguide**

Covalent or molecular compounds contain atoms held together by covalent bonds. These bonds form when the atoms share electrons because they have similar electronegativity values. Covalent compounds are a diverse group of molecules, so there are several

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exceptions to each 'rule'.

## **Covalent or Molecular Compound Properties**

of covalent bonds; • explain the ...

CHEMICAL BONDING AND MOLECULAR  
STRUCTURE 103 When combining atoms  
share three electron pairs as in the case  
of two nitrogen atoms in the N<sub>2</sub>

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molecule and the two carbon atoms in the ethyne molecule, a triple bond is formed. 4.1.3 Lewis Representation of Simple

## **CHEMICAL BONDING AND MOLECULAR STRUCTURE**

A single covalent bond is represented by a small line (–) between the two atoms.

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27. Triple covalent bond : Bond formed due to the sharing of three electron-pairs is called a triple covalent bond or simply a triple bond. 28. Valence electrons : Valence is one of the most important chemical property of the elements.

## **Chemical bonding and molecular**

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- **Covalent Bond—Lewis-Langmuir Concept** When the bond is formed between two or more atoms by mutual contribution and sharing of electrons, it is known as covalent bond. If the combining atoms are same the covalent molecule is known as homoatomic. If they are different, they are known as



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heteroatomic molecule. For Example,

## **Chemical Bonding and Molecular Structure Class 11 Notes ...**

The methane molecule,  $\text{CH}_4$ , can be used to illustrate the procedure for predicting molecular shape. The Lewis structure of this molecule ascribes four bonding electron pairs to the carbon

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atom (Figure 8). These pairs repel one another, and their separation is maximized if they adopt a tetrahedral disposition around the central carbon atom.

## **Chemical bonding - Molecular shapes and VSEPR theory ...**

Filed Under: Chemical Bonding and

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Molecular Structure, Chemistry, Class 11  
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lone pair, single covalent bond, triple  
covalent bond. About Mrs Shilpi Nagpal

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## **Covalent Bond | Chemical Bonding and Molecular Structure ...**

Single and multiple covalent bonds. This is the currently selected item. Metallic bonds. Drawing Lewis diagrams. Predicting bond type (metals vs. nonmetals) Worked example: Lewis diagram of formaldehyde ( $\text{CH}_2\text{O}$ ) ...  
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Chemical bonding and molecular structure ...

### **Single and multiple covalent bonds (article) | Khan Academy**

The structure of diamond. The giant covalent structure of diamond. Carbon has an electronic arrangement of 2,4. In diamond, each carbon shares electrons

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with four other carbon atoms - forming four single bonds. In the diagram some carbon atoms only seem to be forming two bonds (or even one bond), but that's not really the case.

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