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The Clausius-Clapeyron equation: $dP_{i, \text{sat}} = \frac{h_{i, \text{vap}}}{RT^2} dT$ or $\ln P_{i, \text{sat}} = \frac{h_{i, \text{vap}}}{R} \left(\frac{1}{T} - \frac{1}{T_0} \right) + \ln P_{i, \text{sat},0}$
101 kPa [] " # \$ % & ' = () h_{i, \text{vap}} R 1 T
(1 373 [K] * + , - . / so $P_{i, \text{sat}} = 101 \dots$

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thermodynamic equilibrium any
thermodynamic variable for a pure
substance, like pure water, can be
written in terms of any two other
thermodynamic variables, i.e. $p = p(p, T)$
(6.1.1) where the functional relationship
in depends on the substance.

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thermodynamics and present them in a logical, relatively easy to understand manner.

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